



Cambridge International AS & A Level

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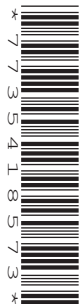


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CHEMISTRY

9701/32

Paper 3 Advanced Practical Skills 2

May/June 2025

2 hours

You must answer on the question paper.

You will need: The materials and apparatus listed in the confidential instructions

INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

INFORMATION

- The total mark for this paper is 40.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.
- Important values, constants and standards are printed in the question paper.
- Notes for use in qualitative analysis are provided in the question paper.

Session

Laboratory

For Examiner's Use

1	
2	
3	
Total	

This document has **12** pages. Any blank pages are indicated.



Quantitative analysis

Read through the whole method before starting any practical work. Where appropriate, prepare a table for your results in the space provided.

Show the precision of the apparatus you used in the data you record.

Show your working and appropriate significant figures in the final answer to **each** step of your calculations.

- 1 Acids donate H^+ in aqueous solution. The number of moles of H^+ donated per mole of acid is the **proticity** of the acid. For example, sulfuric acid is diprotic as it donates two moles of H^+ per mole of acid.

In this experiment you will carry out a titration to determine the proticity of citric acid, $\text{C}_6\text{H}_8\text{O}_7$.

FB 1 is aqueous citric acid containing $7.50 \text{ g dm}^{-3} \text{ C}_6\text{H}_8\text{O}_7$.

FB 2 is aqueous sodium hydroxide containing $4.50 \text{ g dm}^{-3} \text{ NaOH}$.

FB 3 is thymolphthalein indicator.

(a) Method

- Fill the burette with **FB 2**.
- Pipette 25.0 cm^3 of **FB 1** into a conical flask.
- Add a few drops of **FB 3** to the conical flask.
- Perform a rough titration and record your burette readings in the space below.

The rough titre is cm^3 .

- Carry out as many accurate titrations as you think necessary to obtain consistent results.
- Make sure any recorded results show the precision of your practical work.
- Record, in a suitable form in the space below, all your burette readings and the volume of **FB 2** added in each accurate titration.

Results

I	
II	
III	
IV	
V	
VI	
VII	

[7]

- (b) From your accurate titration results, calculate a suitable mean value to use in your calculations. Show clearly how you obtain the mean value.

25.0 cm^3 of **FB 1** required cm^3 of **FB 2**. [1]



(c) Calculations

- (i) Calculate the amount, in mol, of citric acid in 25.0 cm^3 of **FB 1**.
Show your working.

amount of $\text{C}_6\text{H}_8\text{O}_7 = \dots\dots\dots \text{ mol}$ [2]

- (ii) Calculate the amount, in mol, of sodium hydroxide in the volume of **FB 2** in (b).

amount of $\text{NaOH} = \dots\dots\dots \text{ mol}$ [1]

- (iii) Citric acid is triprotic.
Show whether your results in this experiment support this statement.

[1]

- (iv) Complete the equation for the reaction of citric acid, $\text{C}_6\text{H}_8\text{O}_7$, with sodium hydroxide.

$\text{C}_6\text{H}_8\text{O}_7 + \dots\dots \text{NaOH} \rightarrow \dots\dots\dots + \dots\dots\dots$ [1]

- (v) Citric acid does **not** contain a chiral carbon atom.
Draw a possible structural formula of a molecule of citric acid.

[1]

- (d) A student uses a pipette labelled $25.0 \pm 0.06\text{ cm}^3$ to measure **FB 1**.
The student suggests that it is more accurate to measure the volume of **FB 1** with a burette instead of the pipette.

State whether the student's suggestion is correct. Explain your answer.

.....
..... [1]



- (c) **Experiment 2** is the determination of the enthalpy change, ΔH_2 , for the reaction of one mole of solid citric acid with aqueous sodium hydroxide. In this experiment, aqueous sodium hydroxide, **FB 6**, is used in excess.

Method

- Support the second cup in the beaker.
- Add 4.80–5.00 g of **FB 5** to the cup. Record your weighings.
- Measure and record the temperature of **FB 6** in its container.
- Use the 50 cm³ measuring cylinder to transfer 50.0 cm³ of **FB 6** into the cup with **FB 5**.
- Stir the mixture until the maximum temperature is obtained. Record the maximum temperature.
- Calculate and record the mass of **FB 5** used.
- Calculate and record the temperature change.

Results

I	
II	
III	

[3]

(d) Calculations

Calculate the enthalpy change of reaction, ΔH_2 , in kJ mol⁻¹, for the reaction of 1.00 mol of solid citric acid, C₆H₈O₇, with aqueous sodium hydroxide.
Show your working.

$$\Delta H_2 = \begin{array}{c} \text{.....} \\ \text{sign} \end{array} \begin{array}{c} \text{.....} \\ \text{value} \end{array} \text{ kJ mol}^{-1} \quad [2]$$

- (e) Use your values for ΔH_1 and ΔH_2 to calculate the enthalpy change, ΔH_r , in kJ mol⁻¹, for the reaction of 1.00 mol of aqueous citric acid with aqueous sodium hydroxide.

$$\Delta H_r = \begin{array}{c} \text{.....} \\ \text{sign} \end{array} \begin{array}{c} \text{.....} \\ \text{value} \end{array} \text{ kJ mol}^{-1} \quad [1]$$

[Total: 12]

[Turn over]



Qualitative analysis

For each test you should record all your observations in the spaces provided.

Examples of observations include:

- colour changes seen
- the formation of any precipitate and its solubility (where appropriate) in an excess of the reagent added
- the formation of any gas and its identification (where appropriate) by a suitable test.

You should record clearly at what stage in a test an observation is made.

Where no change is observed, you should write 'no change'.

Where reagents are selected for use in a test, the name or correct formula of the element or compound must be given.

If any solution is warmed, a boiling tube must be used. If a solid is heated, a hard-glass test-tube must be used.

Rinse and reuse test-tubes and boiling tubes where possible.

No additional tests should be attempted.



3 (a) Solutions **FB 7** and **FB 8** each contain one cation and one anion. All of the ions are listed in the Qualitative analysis notes.

(i) Carry out the following tests and record your observations in Table 3.1. Use a 1 cm depth of **FB 7** or **FB 8** in a test-tube for each test.

Table 3.1

test	observations	
	FB 7	FB 8
Test 1 Add an equal volume of aqueous potassium iodide, then		
add excess aqueous sodium thiosulfate.		
Test 2 Add a small spatula measure of zinc powder. Leave the mixture to stand.		
Test 3 Add a few drops of aqueous silver nitrate.		
Test 4 Add aqueous ammonia.		

[6]

(ii) Deduce the formula of **FB 8**.

FB 8 is

[1]

(iii) For **FB 8**, state the numbers of **all** the tests from Table 3.1 that involve a redox reaction.

..... [1]

(iv) Give the ionic equation for **one** of the reactions in **Test 4**. Include state symbols.

..... [1]





- (b) (i) Carry out tests to identify the anion in **FB 7**. The anion does **not** contain sulfur. Record your tests and observations in a suitable form in the space below.

You **must** use a boiling tube if any liquid is heated.

[3]

- (ii) Deduce the formula of **FB 7**.

FB 7 is

[1]

[Total: 13]



* 0000800000009 *



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Qualitative analysis notes

1 Reactions of cations

cation	reaction with	
	NaOH(aq)	NH ₃ (aq)
aluminium, Al ³⁺ (aq)	white ppt. soluble in excess	white ppt. insoluble in excess
ammonium, NH ₄ ⁺ (aq)	no ppt. ammonia produced on warming	–
barium, Ba ²⁺ (aq)	faint white ppt. is observed unless [Ba ²⁺ (aq)] is very low	no ppt.
calcium, Ca ²⁺ (aq)	white ppt. unless [Ca ²⁺ (aq)] is very low	no ppt.
chromium(III), Cr ³⁺ (aq)	grey-green ppt. soluble in excess giving dark green solution	grey-green ppt. insoluble in excess
copper(II), Cu ²⁺ (aq)	pale blue ppt. insoluble in excess	pale blue ppt. soluble in excess giving dark blue solution
iron(II), Fe ²⁺ (aq)	green ppt. turning brown on contact with air insoluble in excess	green ppt. turning brown on contact with air insoluble in excess
iron(III), Fe ³⁺ (aq)	red-brown ppt. insoluble in excess	red-brown ppt. insoluble in excess
magnesium, Mg ²⁺ (aq)	white ppt. insoluble in excess	white ppt. insoluble in excess
manganese(II), Mn ²⁺ (aq)	off-white ppt. rapidly turning brown on contact with air insoluble in excess	off-white ppt. rapidly turning brown on contact with air insoluble in excess
zinc, Zn ²⁺ (aq)	white ppt. soluble in excess	white ppt. soluble in excess

2 Reactions of anions

anion	reaction
carbonate, CO ₃ ²⁻	CO ₂ liberated by dilute acids
chloride, Cl ⁻ (aq)	gives white ppt. with Ag ⁺ (aq) (soluble in NH ₃ (aq))
bromide, Br ⁻ (aq)	gives cream/off-white ppt. with Ag ⁺ (aq) (partially soluble in NH ₃ (aq))
iodide, I ⁻ (aq)	gives pale yellow ppt. with Ag ⁺ (aq) (insoluble in NH ₃ (aq))
nitrate, NO ₃ ⁻ (aq)	NH ₃ liberated on heating with OH ⁻ (aq) and Al foil
nitrite, NO ₂ ⁻ (aq)	NH ₃ liberated on heating with OH ⁻ (aq) and Al foil; decolourises acidified aqueous KMnO ₄
sulfate, SO ₄ ²⁻ (aq)	gives white ppt. with Ba ²⁺ (aq) (insoluble in excess dilute strong acids); gives white ppt. with high [Ca ²⁺ (aq)]
sulfite, SO ₃ ²⁻ (aq)	gives white ppt. with Ba ²⁺ (aq) (soluble in excess dilute strong acids); decolourises acidified aqueous KMnO ₄
thiosulfate, S ₂ O ₃ ²⁻ (aq)	gives off-white/pale yellow ppt. slowly with H ⁺





3 Tests for gases

gas	test and test result
ammonia, NH_3	turns damp red litmus paper blue
carbon dioxide, CO_2	gives a white ppt. with limewater
hydrogen, H_2	'pops' with a lighted splint
oxygen, O_2	relights a glowing splint

4 Tests for elements

element	test and test result
iodine, I_2	gives blue-black colour on addition of starch solution

Important values, constants and standards

molar gas constant	$R = 8.31 \text{ J K}^{-1} \text{ mol}^{-1}$
Faraday constant	$F = 9.65 \times 10^4 \text{ C mol}^{-1}$
Avogadro constant	$L = 6.02 \times 10^{23} \text{ mol}^{-1}$
electronic charge	$e = -1.60 \times 10^{-19} \text{ C}$
molar volume of gas	$V_m = 22.4 \text{ dm}^3 \text{ mol}^{-1}$ at s.t.p. (101 kPa and 273 K) $V_m = 24.0 \text{ dm}^3 \text{ mol}^{-1}$ at room conditions
ionic product of water	$K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ (at 298 K (25 °C))
specific heat capacity of water	$c = 4.18 \text{ kJ kg}^{-1} \text{ K}^{-1}$ (4.18 $\text{J g}^{-1} \text{ K}^{-1}$)





The Periodic Table of Elements

Group																			
1	2													13	14	15	16	17	18
		Key																	
		1 Hhydrogen1.0																	

lanthanoids

57	La	58	Ce	59	Pr	60	Nd	61	Pm	62	Sm	63	Eu	64	Gd	65	Tb	66	Dy	67	Ho	68	Er	69	Tm	70	Yb	71	Lu
lanthanum	cerium	praseodymium	neodymium	promethium	samarium	europlum	gadolinium	terbium	dysprosium	holmium	erbium	thulium	ytterbium	lutetium	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—
138.9	140.1	140.9	144.2	—	150.4	152.0	157.3	158.9	162.5	164.9	167.3	168.9	173.1	175.0	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—

actinoids

89	Ac	90	Th	91	Pa	92	U	93	Np	94	Pu	95	Am	96	Cm	97	Bk	98	Cf	99	Es	100	Fm	101	Md	102	No	103	Lr
actinium	thorium	protactinium	uranium	neptunium	plutonium	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—
—	232.0	231.0	238.0	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—

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